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The practice of bartering (trading one thing for another) has existed since the beginning of time. In the past, for example, things like chickens were traded for newspapers. Person A had something that person B wanted, and vice versa. So when person A and person B exchanged items, each had something new. Some chemical reactions are like those made up of spare parts, and the products are new materials. A double substitution reaction is a reaction in which the positive and negative ions of two ionic compounds exchange to form two new compounds. The general shape of a double shift reaction is:  $\text{AB} + \text{CD} \rightarrow \text{AD} + \text{CB}$ . In this reaction,  $\text{A}$  and  $\text{C}$  are positively charged cations, while  $\text{B}$  and  $\text{D}$  are negatively charged anions. Double substitution reactions usually occur between substances in aqueous solution. For a reaction to occur, one of the products is usually a solid precipitate, a gas or a molecular compound such as water. A precipitate forms in a double-substitution reaction when the cations of one reactor combine with the anions of the other reactor to form an insoluble ionic compound. When mixing aqueous solutions of potassium iodide and lead (II) nitrate, the following reaction occurs:  $2\text{KI} + \text{Pb}(\text{NO}_3)_2 \rightarrow \text{PbI}_2 + 2\text{KNO}_3$ . Lead iodide precipitate formation. (CC BY-NC; CK-12) Some double-substitution reactions produce a gaseous product which then escapes from the solution and escapes into the air. When solutions of sodium sulphide and hydrochloric acid are mixed, the reaction products are aqueous sodium chloride and gaseous hydrochloric acid.  $\text{Na}_2\text{S} + 2\text{HCl} \rightarrow 2\text{NaCl} + \text{H}_2\text{S}$ . Molecular as one of its products. Many examples in this category are water-producing reactions. When aqueous hydrochloric acid is reacted with aqueous sodium hydroxide, the products are aqueous sodium chloride and water:  $\text{HCl} + \text{NaOH} \rightarrow \text{NaCl} + \text{H}_2\text{O}$ . Example (PageIndex{1}) Write a complete and balanced chemical equation for the following double substitution reactions. A product is indicated as a guide.  $\text{NaCN} + \text{HBr} \rightarrow \text{NH}_4\text{SCN} + \text{Ba}(\text{NO}_3)_2$  (a precipitate of sulphate barium forms) Solution Step 1: Plan the problem. In A, the production of a gas pushes the reaction. In B, the production of a precipitate leads the reaction. In both cases, use the ionic accusations of both reagents to build the correct formulas of the products. Step 2: Fix. A. The cations of both reagents are  $\text{Na}^+$  and  $\text{CN}^-$  charged ions, while the anion are  $\text{H}^+$  and  $\text{Br}^-$  charged ions. After the exchange of the partners, the balanced equation is:  $\text{NaCN} + \text{HBr} \rightarrow \text{NaBr} + \text{HCN}$ . B. This must be taken into account when exchange partners and writing new formulas. So the equation is balanced.  $\text{Na}_2\text{CO}_3 + 2\text{HCl} \rightarrow 2\text{NaCl} + \text{CO}_2 + \text{H}_2\text{O}$ . Double replacement reactions can form precipitates, gases or molecular compounds. Occasionally, a reaction will produce both a gas and a molecular compound. Contributors and attributions Foundation CK-12 by Sharon Bewick, Richard Parsons, Therese Forsythe, Shonna Robinson and Jean Dupon. The questions published on the site are exclusively generated by the user, doubting has no ownership or control over the nature and content of such questions. The doubt is not responsible for any discrepancies regarding the duplexity of the content regarding those questions. questions.

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